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Naming Molecular Compounds

How are the chemical formula and name of a molecular compound related?

Why?

When you began chemistry class this year, you probably already know that the chemical formula for carbon dioxide was CO_2 . Today you will find out why CO_2 is named that way. Naming chemical compounds correctly is of paramount importance. The slight difference between the names carbon monoxide (CO , a poisonous, deadly gas) and carbon dioxide (CO_2 , a greenhouse gas that we exhale when we breathe out) can be the difference between life and death! In this activity you will learn the naming system for molecular compounds.

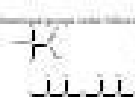
Model 1 – Molecular Compounds


Molecular Formula	Number of Atoms of First Element	Number of Atoms of Second Element	Name of Compound
ClF			Chlorine monofluoride
ClF_5	1	5	Chlorine pentafluoride
CO			Carbon monoxide
CO_2			Carbon dioxide
Cl_2O			Dichlorine monoxide
PCl_5			Phosphorus pentachloride
N_2O_5			Dinitrogen pentoxide


- Fill in the table to indicate the number of atoms of each type in the molecular formula.
- Examine the molecular formulas given in Model 1 for various molecular compounds.
 - How many different *elements* are present in each compound shown?
 - Do the compounds combine metals with metals, metals with nonmetals, or nonmetals with nonmetals?
 - Based on your answer to *b*, what type of bonding must be involved in molecular compounds?
- Find all of the compounds in Model 1 that have chlorine and fluorine in them. Explain why the name "chlorine fluoride" is not sufficient to identify a specific compound.
- Assuming that the name of the compound gives a clue to its molecular formula, predict how many atoms each of these prefixes indicates, and provide two examples.
 - mono-
 - di-
 - penta-

Electron's General Chemistry answer Worksheet

1. Identify the compound's name and write the corresponding formula.

a.  _____

b.  _____

c.  _____

2. Write the following names and formulas and use the space to draw the structure.

a. H_2 _____

b. O_2 _____

c. N_2 _____

d. C_2 _____

Naming Chemical Compounds - Answers

Name the following *ionic* compounds:

- NaBr **sodium bromide**
- CaO **calcium oxide**
- Li_2S **lithium sulfide**
- MgBr_2 **magnesium bromide**
- $\text{Be}(\text{OH})_2$ **beryllium hydroxide**

Write the formulas for the following *ionic* compounds:

- potassium iodide **KI**
- magnesium oxide **MgO**
- aluminum chloride **AlCl_3**
- sodium nitrate **NaNO_3**
- calcium carbonate **CaCO_3**
- lithium sulfate **Li_2SO_4**
- beryllium phosphide **Be_3P_2**
- magnesium hydroxide **$\text{Mg}(\text{OH})_2$**
- sodium phosphate **Na_3PO_4**
- aluminum carbonate **$\text{Al}_2(\text{CO}_3)_3$**
- calcium chloride **CaCl_2**
- sodium cyanide **NaCN**
- aluminum oxide **Al_2O_3**
- magnesium acetate **$\text{Mg}(\text{C}_2\text{H}_3\text{O}_2)_2$**
- ammonium chloride **NH_4Cl**

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